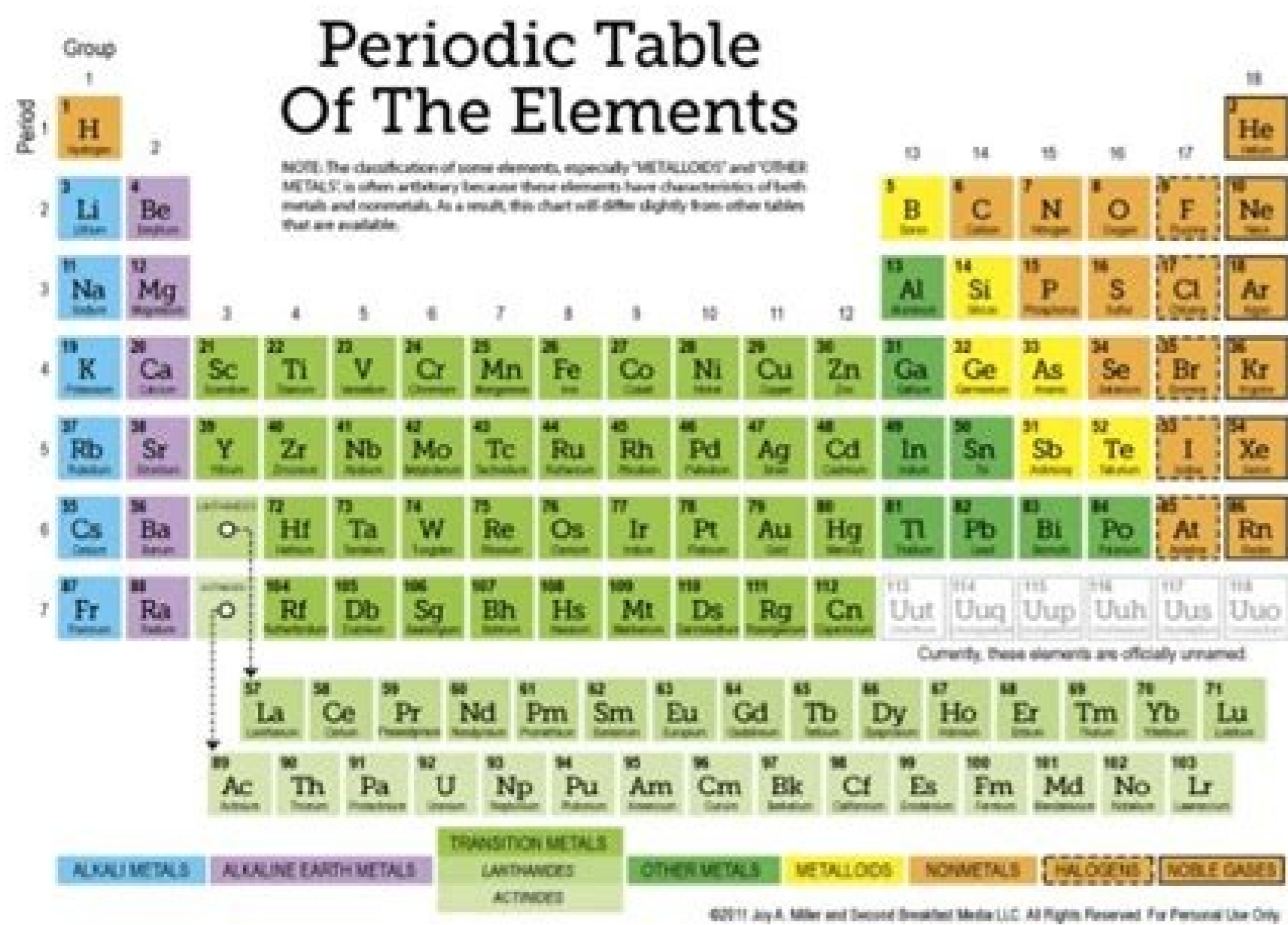


I'm not robot!



Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Atomic Structure – Ch. 3

### PART A – SUBATOMIC PARTICLES

The table below contains information about several elements. In each case, enough information has been provided for you to fill in the blanks. Assume all atoms are neutral.

Isotope Name	Nuclear Symbol	Atomic Number	Mass Number	# of Protons	# of Electrons	# of Neutrons
1. calcium-40						
2.		12	24			
3.				1		2
4.	$^{197}_{79}\text{Au}$					
5.					26	30
6.			201	80		
7.		17				18

### PART B – AVERAGE ATOMIC MASS

- Calculate the average atomic mass for neon if its abundance in nature is 90.5% neon-20, 0.3% neon-21, and 9.2% neon-22.
- Calculate the average atomic mass of silver if 13 out of 25 atoms are silver-107 and 12 out of 25 atoms are silver-109.
- Distinguish between mass number, relative atomic mass, and average atomic mass.

NAME \_\_\_\_\_

## Periodic Table Worksheet 1

There are FOUR parts that are displayed for every element on every Periodic Table. Use this labeled picture and the Periodic Table as your guide to complete the worksheet.

**DIRECTIONS:** You are given ONE of the FOUR parts of the element's square – find the element using the given information and fill in the three missing parts of the square.

C	Mn	Pt	1	31	
					Vanadium
24		5	Fe		52
	Helium			Cobalt	
32.065	Re	83.798	37	Ce	Fermium
				87	Hf
Iridium	200.59	58.693	Silicon		

